
Investigation of Metal-Oligonucleotide Complexes by Nanoelectrospray Tandem Mass Spectrometry in the Positive Mode

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The formation and fragmentation of multiply metal-coordinated oligonucleotides was studied by nanoelectrospray tandem mass spectrometry in the positive ion mode. Fundamental aspects of the gas-phase behavior of metal-oligonucleotide complexes are revealed. The addition of transition metal ions, such as iron(II), iron(III), and zinc(II), leads to very stable metal-oligonucleotide complexes which show heavily altered fragmentation patterns in contrast to uncomplexed oligonucleotides. The site of metal ion complexation was located by collision-induced dissociation (CID) experiments. It was found that all three metal ions investigated predominantly coordinate to the central phosphate groups of the oligonucleotides. Furthermore, it is demonstrated that the fragmentation of such complexes depends highly upon the metal ion complexed as well as on the sequence of the nucleobases in the oligonucleotide. (*J Am Soc Mass Spectrom* 2005, 16, 370–378) © 2004 American Society for Mass Spectrometry

Metal ions play a central role in cellular processes. Besides their function as cofactors of proteins, especially in replication, transcription, and translation, they induce folding and cross-linking of nucleic acids and result in a stabilization of these structures. Such interaction can be highly specific, like the linkage of two guanines by cisplatin complexes, or unspecific, like backbone charge neutralization on a DNA double strand [1, 2]. Likewise, metal ions are involved in the stabilization of RNA, ribozymes, and telomers [3–5]. In medicine, metal ions find broad therapeutic application on DNA, as cisplatin complexes and related drugs form DNA-adducts, which act as toxic agents in tumor tissue [6]. These drugs are widely applied even though the mechanism on the molecular level is not fully elucidated yet [7]. Furthermore, there are approaches to use metal ions as markers for improved diagnostic applications [8–14]. The application of metal ions may even hold more potential, which can be taken advantage of once the interactions of metal ions with nucleic acid structures are better understood. Several NMR studies have been published with the aim of elucidating the metal ion binding pattern [15–17]. Although several approaches have been made to determine the coordination sites, no general rule on how metal ions bind to nucleic acid structures could be established yet.

The interaction of metal ions with nucleic acids has

also been investigated using mass spectrometry. In an early study, Christian et al. describe the influence of iron on oligonucleotides as matrix dependent [18]. Another study describes the influence of Mg^{2+} , Na^+ , and UO_2^{2+} on oligonucleotides and demonstrates that the phosphodiester groups are considered the most probable binding sites. Additionally, it is reported that these metal ions bind preferentially to the central thymine region of the analyzed octa- and dodecamer [14].

By exploring oligonucleotides with their phosphate groups completely occupied by alkali or earthalkali ions, Gross and coworkers observed that the product ion spectra of such metal complexed oligonucleotides differ significantly from the spectra of their uncomplexed analogs [19]. For the first time the fragmentation mechanism of such complexes was investigated and it was postulated that dissociation of the complexes is no longer dependent upon the proton affinity of the nucleobases, but rather on the type of the nucleobase. Also, it was remarked that metal complexed oligonucleotides show increased stability in gas phase. Another study located sodium ions at the central phosphate groups and found the position of the sodium ions to be dependent upon the number of phosphate groups [20]. Most recently, a detailed study on the interaction of deoxydinucleotides with alkali metal ions has been published by Xiang et al., demonstrating that the fragmentation of the complexed molecules depends upon the sequence, the interaction of nucleobase and metal ion, and the steric hindrance of the nucleobases [21]. A comprehensive study on iron-oligonucleotide complexes using matrix-assisted laser desorption/ionization mass spectrometry (MALDI) has been published by

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Hettich [22, 23]. Generally, published work on mass spectrometry of metal-oligonucleotide complexes focuses on electrospray ionization (ESI) and MALDI in the negative ion mode. Analyzing oligonucleotides in the positive ion mode is not common, though data have been reported previously [24–29].

In this article, we report detailed data on metal-oligonucleotide complexes analyzed in the positive ion mode, since charge compensation due to the presence of positively charged metal ions greatly decreases the number of detectable complexes upon analysis in the negative ion mode.

Information on binding sites and on the altered dissociation behavior of oligonucleotides in the presence of metal ions is given, thereby furnishing an additional piece in the fundamental research on the influence of metal ions on nucleic acid structures. In addition, the influence of the metal ions on the fragmentation behavior of oligonucleotides shall be clarified to characterize specific patterns in terms of future applications, e.g., for directed fragmentation of oligonucleotides, potentially providing an aid in sequencing nucleic acid structures by mass spectrometry. In the present study, the behavior of metal-oligonucleotide complexes upon CID is elucidated in-depth with the aim to exactly determine the binding sites of the metal ions within the molecules based on the structural information gained by the product ion spectra.

Experimental

Sample Preparation

Oligonucleotides used were unmodified, single stranded oligodeoxynucleotides of the sequences TTCAT, TTGAT, TTGCT, and TTGGT. They were obtained from Microsynth (Balgach, Switzerland) and used without further purification. High purity water was purchased from Fluka (Buchs, Switzerland). Iron(II)-chloride tetrahydrate, iron(III)-chloride hexahydrate, and zinc-chloride, all in puriss p.a. quality, were also purchased from Fluka. Aqueous solutions of oligonucleotides and metal salts were prepared with concentrations of 10–40 pmol/ μ l and 200 pmol/ μ l, respectively.

Mass Spectrometry

All experiments were performed on an Applied Biosystems/MDS Sciex QSTAR Pulsar hybrid quadrupole-time-of-flight mass spectrometer (Sciex, Concord, Canada), equipped with a nanoelectrospray ion source (Proxeon Biosystems, Odense, Denmark). Oligonucleotides were analyzed in the positive ion mode with a potential of 800 to 1000 V applied to the nanospray needle. Nitrogen was used as the curtain gas. Tandem mass spectrometric experiments were performed with the precursor ions selected within a window of ± 1.5 m/z units. Collision-induced dissociation was per-

formed with collision energies in the range of 10–35 eV using nitrogen as the collision gas. Calibration of the instrument was done with cesium iodide and reserpine (Fluka). The Applied Biosystems Analyst QS software package was used for data processing.

Results and Discussion

Investigation of metal-ion oligonucleotide complexes by electrospray tandem mass spectrometry was performed with a selection of four pentanucleotides containing thymine, which exhibits a relatively low proton affinity, as the terminal nucleobases. The other three nucleobases served as exchangeable units in the central part of the sequences. Relatively short oligonucleotides with asymmetric base sequence are well suited as model compounds, as their analysis results in spectra without unnecessary complexity, while they still provide the desired information regarding the fragmentation pattern.

Figure 1 shows the complexes of TTCAT in the presence of $ZnCl_2$. Besides the peak of the uncomplexed, doubly protonated oligonucleotide, peaks corresponding to zinc-oligonucleotide complexes are present in the spectrum. Up to three Zn^{2+} -ions are bound to the oligonucleotide and despite this varying number of metal ions, all complexes are doubly charged. Obviously, the charge state of the complexes is independent of the number of attached zinc ions. The characteristic isotopic distribution of zinc is accurately reflected by the spectrum and the isotopic patterns observed for the various complexes are in excellent agreement with the corresponding calculated distributions.

Metal-oligonucleotide complexes were selected as pre-

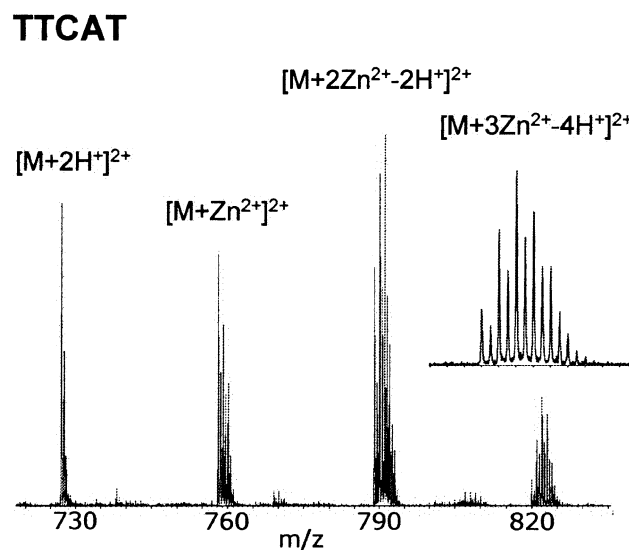
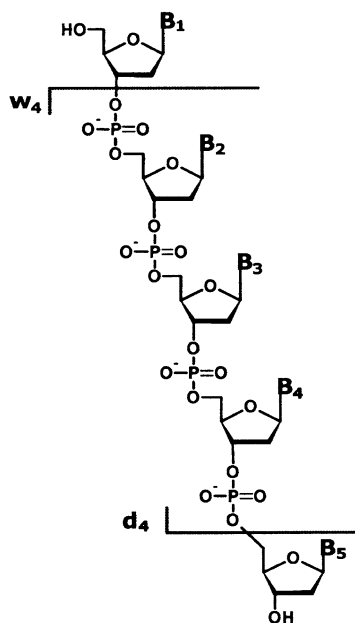


Figure 1. Section of the nanospray mass spectrum of a solution containing the oligonucleotide TTCAT and $ZnCl_2$. The multiple metal adducts can clearly be identified. The inset shows the enlarged isotopic distribution of the triple zinc complex.

cursor ions for CID experiments. The presence of metal ions leads to a considerable change of the fragmentation behavior. The fragmentation of Fe^{3+} -, $[\text{Fe}^{2+}]_3$ -, and $[\text{Zn}^{2+}]_3$ -pentamer complexes was investigated and all product ion spectra showed significant alterations of the fragment ion pattern compared to the product ion spectra of the uncomplexed pentamers. Although the ion patterns largely depend on the type of complex, they do reflect some general properties inherent to all metal-pentamer complexes. It was observed that the abundant peaks of the w_2 -ion, found in the product ion spectra of the uncomplexed oligonucleotides, are generally missing in the product ion spectra of the metal-oligonucleotide complexes. Also, suppression of the a - and $(a\text{-}B)$ -ion series is increased for the complexes, when compared to the uncomplexed oligonucleotides.

The $(a_4\text{-}B_4)$ -ions show a noticeably decreased abundance in the product ion spectra of the Fe^{3+} -pentamer complexes and are completely missing in the product ion spectra of $[\text{Fe}^{2+}]_3$ - and $[\text{Zn}^{2+}]_3$ -pentamer complexes. Generally, an increase of metal-complexed $M\text{-}B$ -, w_4 -, and $(w_4\text{-}B)$ -ions in the product ion spectra of metal-ion oligonucleotide complexes is observed, thus, indicating a stabilization of the phosphate backbone by metal ions. This finding is confirmed by the presence of the metal-complexed core fragment $[w_4\text{-}d_4]$ (Scheme 1), which is found exclusively in the product ion spectra of the metal-ion oligonucleotide complexes. The presence of the $[w_4\text{-}d_4]$ -ion suggests that the stabilizing effect originates from the center of the molecule. This change from the product ion spectra of the uncomplexed oligonucleotide to the product ion spectra of the metal-complexed oligonucleotide is illustrated in Figure 2.

Metal-coordinated fragment ions originating from either a $[\text{Fe}^{2+}]_3$ - or a $[\text{Zn}^{2+}]_3$ -pentamer complex always



Scheme 1

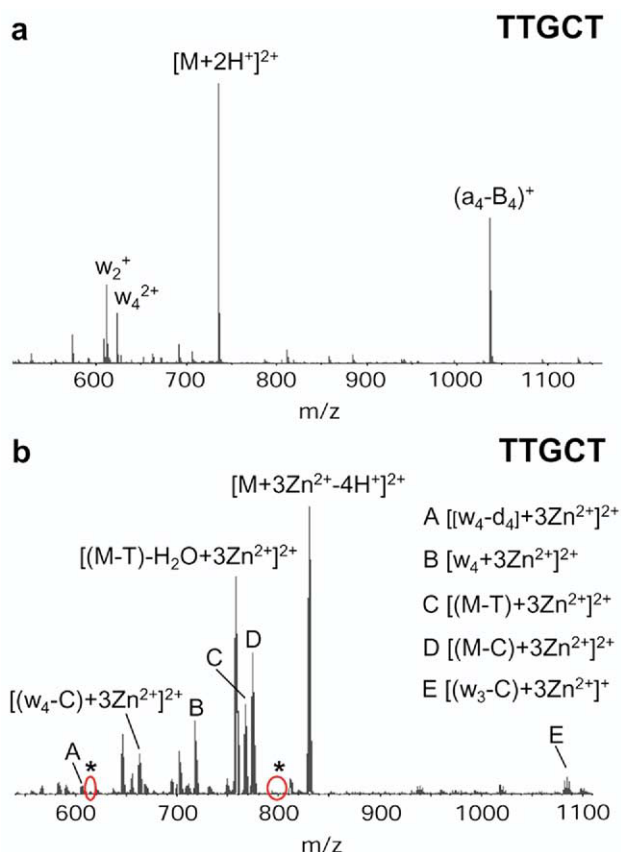


Figure 2. Comparison of the product ion spectra of $[\text{TTGCT} + 2\text{H}^+]^{2+}$ (a) and $[\text{TTGCT} + 3\text{Zn}^{2+} - 4\text{H}^+]^{2+}$ (b). The intensive peak of the w_2 -fragment visible in the spectrum of the uncomplexed oligonucleotide is not detected in the spectrum of the $[\text{Zn}^{2+}]_3$ -complex, neither in the uncomplexed nor in the metal-complexed form (asterisks). On the other hand, the abundant peaks of the $M\text{-}B$ -ions of the $[\text{Zn}^{2+}]_3$ -complex in the spectrum of the uncomplexed oligonucleotide lack, whereas abundant peaks of the w_4 -ion are visible in both spectra. All ions labeled in the spectrum of $[\text{Zn}^{2+}]_3\text{-TTGCT}$ (b) bear all three Zn^{2+} -ions, as proven by the characteristic isotopic pattern.

bear three metal ions, identical to the number of metal ions coordinated to their precursor. A few low-mass fragments, which do not show any metal-coordination at all, can be found as well. Yet, no fragment ions descending from a triply complexed precursor ion coordinating only one or two metal ions have been identified, thus proving that the bonds formed between metal ions and oligonucleotide are very strong. This is demonstrated by the spectrum in Figure 3a. By forming such strong complexes, the metal ions effect a stabilization, particularly of the phosphate-backbone, changing the fragmentation pattern of the complexes compared to the fragmentation pattern of bare oligonucleotides, as described above. Such an effect seems to be a characteristic property of the transition metal investigated, and does not apply to metal-oligonucleotide complexes in general, as a comparative CID experiment with a $[\text{Na}^+]_6$ -pentamer complex impressively revealed. In contrast to the dissociation of transition metal ion

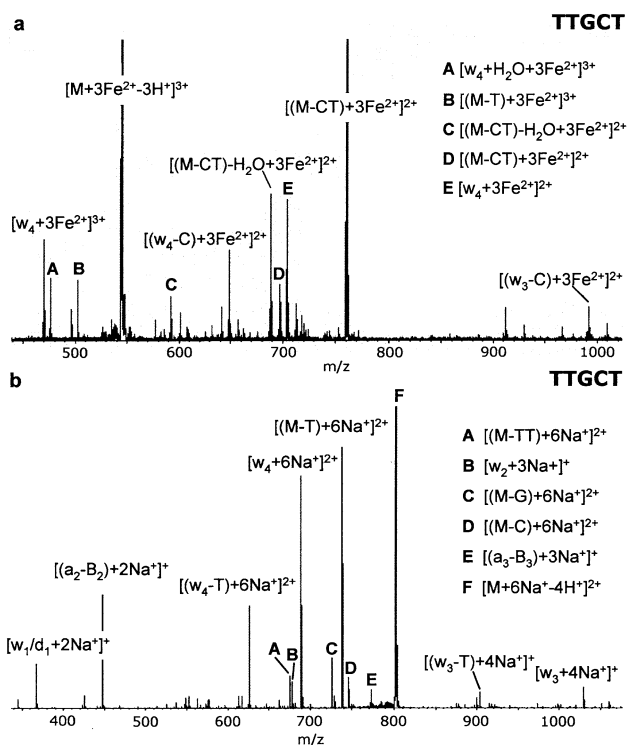


Figure 3. (a) CID spectrum of $[TTGCT + 3Fe^{2+} - 3H^+]^{3+}$. Every fragment coordinates exactly three Fe^{2+} -ions, identical to the number of iron ions complexed in the precursor ion. Consequently, very strong bonds are formed between oligonucleotide and metal ions, as metal ions are not released upon CID. This behavior is a characteristic property of the metal-oligonucleotide complexes investigated in this study. (b) CID spectrum of $[TTGCT + 6Na^+ - 4H^+]^{2+}$. In contrast to spectrum in Figure 3a, the number of Na^+ -adducts corresponds to the length of the respective fragment but is independent of the precursor ion. Consequently, there is no stabilizing effect and the fragmentation pattern resembles rather the one of an uncomplexed oligonucleotide.

complexes, fragments with a varying number of metal ions are generated. The number of sodium ions on a fragment ion is chiefly controlled by the length of the fragment, but independent of the total number of metal ions coordinated to its precursor ion (Figure 3b). Sodium is not able to form nearly as strong bonds as the three transition metals investigated. Hence, there is no stabilizing effect and the fragmentation pattern of a sodium-oligonucleotide complex resembles more the one of an uncomplexed oligonucleotide.

The Effect of Fe^{3+} , Fe^{2+} , and Zn^{2+} on the Fragmentation of Four Model Pentamers and Their Localization by Analysis of Given Fragments

Studying the behavior of metal-oligonucleotide complexes in the positive ion mode by tandem mass spectrometry is a step into a hardly discovered area. Interpretation of analytical data is strongly complicated by the large diversity of complexes generated. The diversity due to the combination of different metal ions with oligonucleotides is further increased by the varying

number of coordinated metal ions and resulting charge states. Localization of metal ions within the oligonucleotides is performed by identifying the shortest metal-complexed fragment ions from the 5'- and the 3'-end. The overlapping portion of the two oppositely directed partial sequences is the metal-binding part of the oligonucleotide. This strategy of localization is supported by the few short uncomplexed fragments and the remaining longer metal-complexed fragment ions.

Dissociation of Uncomplexed Pentanucleotides $[M + 2H]^{2+}$

Protonation of the nucleobase is the key step initiating base loss and subsequent backbone dissociation of deprotonated oligodeoxynucleotides [30, 31] and there is evidence that the same mechanism applies to protonated oligodeoxynucleotides as well. Although the corresponding product ion spectra are not entirely identical, dissociation of both positively and negatively charged precursor ions results in characteristic w- and (a-B)-ions [32] of various length as the most abundant fragment ions. This observation is in agreement with other studies on protonated oligonucleotides [26–28].

The most abundant peaks in the spectra of uncomplexed oligonucleotides arise from w_2 -, (a_4 - B_4)- and w_4 -ions. As illustrated in Figure 4, the dissociation of the molecule occurs between all nucleotides. However, the abundances of the dominant fragment ions differ between the spectra of the four oligonucleotides studied, as shown by the relative abundance of the (a_4 - B_4)-ion in the spectra of TTCAT, TTGAT, TTGCT, and TTGGT. In the spectra of TTGGT and TTGCT, the

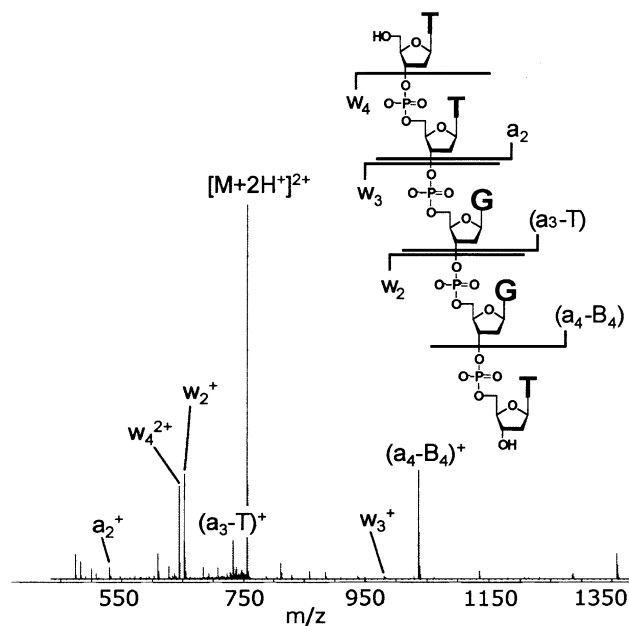


Figure 4. Product ion spectrum of $[TTGGT + 2H^+]^{2+}$. The most intensive peaks arise from the w_2 -, w_4 -, and (a_4 - B_4)-fragment. Upon CID the oligonucleotide dissociates between all nucleotides, giving fragments of varying lengths.

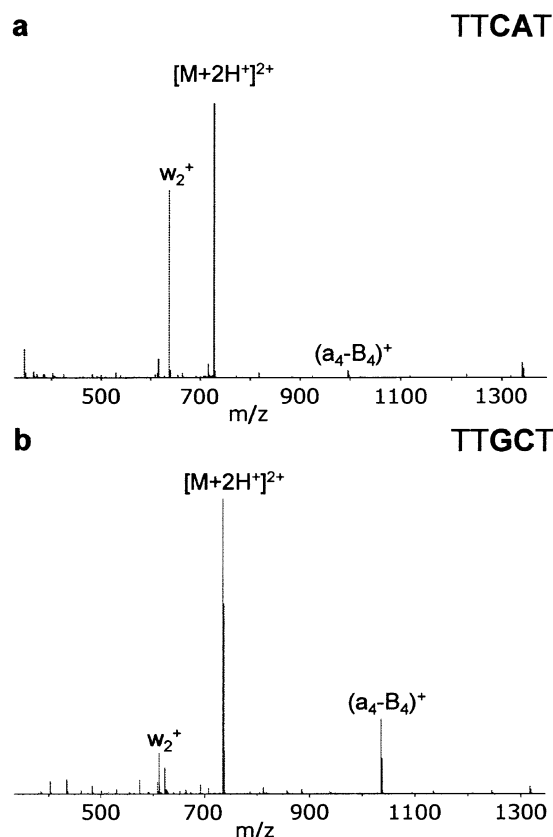


Figure 5. Sequence dependence of the fragmentation on the example of two uncomplexed oligonucleotides. While the oligonucleotide of the sequence TTCAT shows the w_2 -ion with high and the $(a_4-B_4)^+$ -ion only with significantly lower abundance than w_2 (a), the opposite is true for the oligonucleotide TTGCT (b).

(a_4-B_4) -ion occurs with very high abundance, which even exceeds the generally abundant w_2 -ion. Opposite behavior was found for TTGAT and TTCAT. Here, the w_2 -ion shows the highest abundance and the (a_4-B_4) -ion appears only with average intensity comparable to other fragments (Figure 5). The most evident difference between the two pairs of pentanucleotides, which exhibit such diverse fragmentation behavior, is the presence of adenine within TTGAT and TTCAT and its absence in TTGGT and TTGCT. This observation strongly implies that the base sequence of the oligonucleotides does affect the fragmentation.

Dissociation of $[M + Fe^{3+} - H^+]^{2+}$

By adding $FeCl_3$ to the oligonucleotides in solution, iron-oligonucleotide complexes are generated. Complexes of pentanucleotides with one or two Fe^{3+} -ions are observed. Generally, complexes bearing a single Fe^{3+} -ion exhibit higher abundance and have been selected for further investigation. The product ion spectra of the four complexes investigated show the metal-complexed w_4 -ion with highest abundance. Metal-complexed (w_4-B) , M-B and the core fragment $[w_4-d_4]$ lead to abundant peaks as well, whereas the

w_2 -ion, which is very abundant in the product ion spectra of the bare oligonucleotides, is not detected any more, either in metal-complexed, or in uncomplexed form. Also, the abundance of the dominant (a_4-B_4) -ions observed in the spectra of uncomplexed TTGCT and TTGGT has remarkably decreased in the product ion spectra of the Fe^{3+} -complexed oligonucleotides. In their Fe^{3+} -complexed form, the (a_4-B_4) -ions appear to a reduced amount only. Besides the metal-complexed fragments, the uncomplexed a_2 - and z_2 -ions, a_1/z_1^- , c_1/x_1^- , and nucleoside-like ions with m/z 192, 207, 216, and 232, whose structures have been proposed previously [27, 28], are detected as well. The presence of the iron-containing $[w_4-d_4]$ core fragment (Scheme 1) in the spectra of all four pentamers suggests that the Fe^{3+} -ion has a stabilizing effect on the phosphate backbone. This is reflected by the lack of short metal-adducted fragments and the frequent loss of nucleobases. Thus, the presence of the Fe^{3+} -ion clearly alters the fragmentation pathway and enhances the formation of characteristic metal-containing fragments, as demonstrated by the product ion spectrum of $[TTGAT + Fe^{3+} - H^+]^{2+}$ in Figure 6.

Identification of the shortest metal-complexed fragment ions from the 5'- and the 3'-end enables the localization of the Fe^{3+} -ion within the oligonucleotide. Simultaneous occurrence of the metal-complexed fragments $(w_3-B_3)/(w_3-B_4)$, and (a_3-B_3) within the product ion spectra of all four pentamers restricts the position of the Fe^{3+} -ion to the third nucleotide from the 3'-end (Figure 7). Since the nonterminal sugar residues do not provide any suitable binding site for metal ions, the Fe^{3+} -ion must coordinate either to nucleobase B_3 or to the adjacent phosphate group. Binding of the Fe^{3+} -ion to the third nucleobase B_3 can clearly be excluded, as the metal-complexed (a_3-B_3) -ions are present in the spectra of all four Fe^{3+} -pentamer complexes studied. Further evidence is given by detection of the complementary metal-complexed (w_3-B_3) -ions, confirming that nucleobase B_3 is not involved in complexation of the

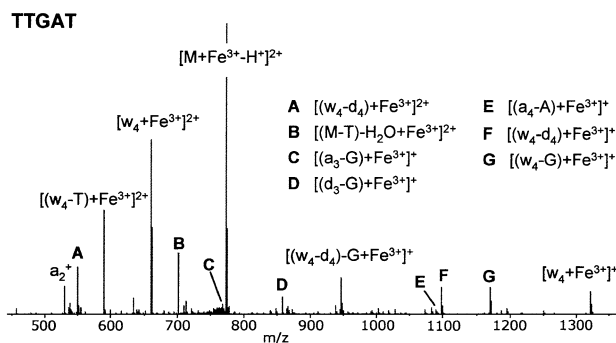


Figure 6. CID spectrum of the complex $[TTGAT + Fe^{3+} - H^+]^{2+}$. All intensive peaks correspond to fragments with a minimum length of three nucleoside units and bear the Fe^{3+} -ion. The only exception is the fragment a_2 which is shorter and also lacks the metal-ion. Specific for the metal-oligonucleotide complex is further the core fragment $[w_4-d_4]$.

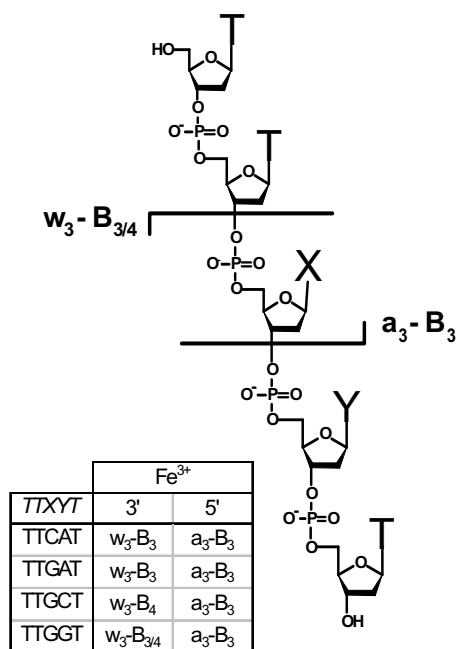


Figure 7. Localization of the Fe³⁺-ion within the four pentamers. The (a₃-B₃)-ions are the shortest 5'-fragments and the w₃-B_{3/4}-ions are the shortest 3'-fragments coordinating the Fe³⁺-ion. Consequently the Fe³⁺-ion must be located in the center of the molecule, on the second phosphate group from the 5'-end.

Fe³⁺-ion. Based on these observations we can conclude that the metal ion must be associated with the phosphate group and consequently, for the pentamers studied, we can accurately locate the Fe³⁺-ion at the second phosphate group from the 5'-end. Also, this finding is in accordance with the observation of uncomplexed a₂- and z₂-fragments (Figure 7).

The sequence specific difference between the pentanucleotides TTCAT/TTGAT and TTGGT/TTGCT, observed in the product ion spectra of the bare oligonucleotides, does not appear if the oligonucleotides are coordinated with a Fe³⁺-ion. Here, the w₄-ion is of highest abundance within the spectra of all four Fe³⁺-oligonucleotide complexes and the abundant metal-complexed (w₄-B)-, (M-B)-, and [w₄-d₄]-ions do not give any evidence for a sequence specific fragmentation behavior. Thus, the Fe³⁺-ion coordinating to the center of the molecule obviously suspends the influence of the nucleobase sequence.

Dissociation of [M + 3Fe²⁺ - 3H⁺]³⁺

Investigating the product ion spectra of triply Fe²⁺-coordinated pentamers reveals three remarkable facts. First, all fragment ions appear exclusively in the metal-complexed form, which is confirmed by the calculated molecular weights and by the characteristic isotopic patterns of iron-adducted fragments. Second, the complementary fragments, e.g., a₂ or (a₂-B₂) as counterparts to (w₃-B), are absent. In the range below m/z 400, peaks corresponding to the released nucleobases and the

uncomplexed nucleoside-like ions with m/z 192 and m/z 232 are visible [27, 28]. The third, rather unexpected observation is the complete lack of any 5'-fragments in the product ion spectra of [M + 3Fe²⁺ - 3H⁺]³⁺. The disappearance of the (a₄-B₄)-fragment ion, which was present in the product ion spectra of both the uncomplexed and the Fe³⁺-complexed pentamers, is remarkable. This behavior can most likely be assigned to a fundamental change in the fragmentation pathway caused by the complexation with metal ions.

The absence of uncomplexed counterions renders the localization of the binding sites of the three Fe²⁺-ions more difficult than the localization of the single Fe³⁺-ion demonstrated previously. For the four pentamers investigated, the shortest identified metal-coordinated fragments are the (w₃-B)-ions, thus limiting the binding sites of the metal ions to the three 3'-terminal nucleotide units. A further positional constraint is provided by the observation of the [w₄-d₄] core fragment, which excludes both the 3'- and the 5'-end as binding sites. Consequently, potential binding sites for the three Fe²⁺-ions are restricted to the third and fourth nucleotide and the adjacent phosphate group, as shown in Figure 8.

As found previously for the Fe³⁺-adducted pentamers, Fe²⁺-ions are also primarily located in the central region of the pentamers, where they stabilize the phosphate-backbone. The frequent loss of nucleobases, indicated by the presence of M-B-, (w₄-B)-, and (w₃-B)-

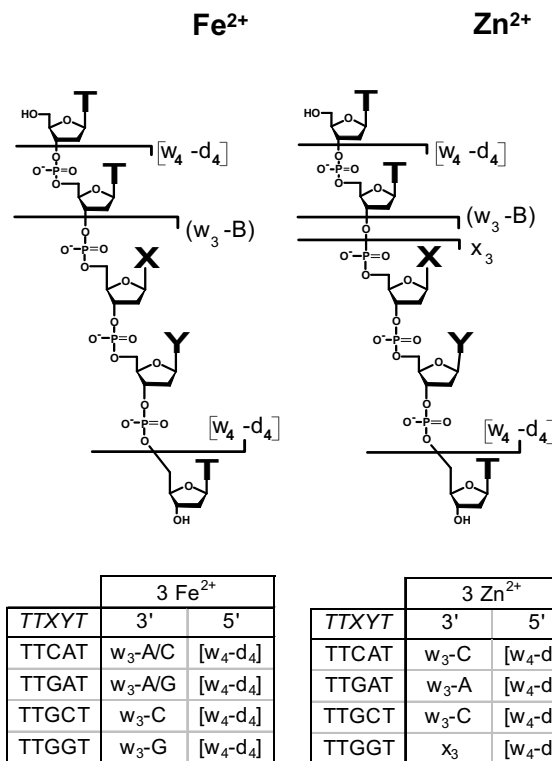


Figure 8. [TXXYT + 3M²⁺]. Localization of the three Fe²⁺/Zn²⁺-ions within the four pentamers. The shortest 3'- and 5'-fragments which coordinate the metal ions are listed and marked with a bar. For details see text.

fragments, strongly suggests that the primary binding sites of the metal ions are the phosphate groups of the backbone and not the nucleobases. However, experiments with $[\text{Fe}^{2+}]_4$ -pentamer complexes and the detection of a fivefold Fe^{2+} -complexed pentamer (data not shown) suggest that the nucleobases can as well be involved in the binding of metal ions, since coordination of more than one doubly-charged metal ion per phosphate group seems utmost unlikely due to the resulting coulombic interactions. Therefore, we conclude that all available phosphate groups are predominantly occupied by Fe^{2+} -ions and if additional binding sites are required, the nucleobases can serve as binding sites too.

Different behavior of TTCAT/TTGAT on the one side and TTGCT/TTGGT on the other side, which was observed previously for uncomplexed oligonucleotides, is also found for the $[\text{Fe}^{2+}]_3$ -pentamer complexes. Here, the difference is reflected by the abundance of metal-complexed w_4 - and M-B-ions. Within the sequences TTCAT and TTGAT the generation of w_4 -ions is strongly favored, followed by M-T-ions. Surprisingly, M-C, M-G, as well as M-A-ions are missing. In contrast to these findings, dissociation of TTGCT and TTGGT primarily generates metal-complexed M-C- and M-G-ions as the main fragments. Repetitive loss of two nucleobases, resulting in M-C-T- and M-G-T-ions, is registered to a noticeable amount too. These results suggest that the presence of adenine has a destabilizing effect within the pentamers, despite the stabilizing influence of the metal ions coordinated.

*Dissociation of $[M + 3\text{Zn}^{2+} - 4\text{H}^+]^{2+}$
and $[M + 3\text{Zn}^{2+} - 3\text{H}^+]^{3+}$*

The spectra of $[\text{Zn}^{2+}]_3$ -pentamer complexes show abundant peaks of metal-coordinated w_4 - and M-B-ions. To a lesser extent, metal-coordinated $(w_4\text{-B})^-$, $[w_4\text{-d}_4]^-$, and $(w_3\text{-B})^-$ -ions are detected as well. Peak assignment may be hindered by the formation of isobaric 5'- and 3'-ions. This problem was also encountered in the case of zinc-adducted pentanucleotides, as abundant peaks corresponding to metal-coordinated c_4 -/ x_4 -ions were found in the spectra of $[\text{Zn}^{2+}]_3$ -TTGAT and $[\text{Zn}^{2+}]_3$ -TTGGT. Additionally, a_4 -/ z_4 -ions were observed in the case of $[\text{Zn}^{2+}]_3$ -TTGGT. However, the absence of any unambiguously assigned 5'-fragment ion and the presence of the x_3 -fragment ion in the spectrum of $[\text{Zn}^{2+}]_3$ -TTGGT suggests that the fragment ions in question rather derive from the 3'-end, thus, are the x_4 - and z_4 -ions.

Furthermore, short uncomplexed fragment ions are observed, whereof the one with m/z 225 can be assigned to a_1 , being the counterion to the w_4 -fragment. Peaks at m/z 192, 216, and 232 correspond to nucleoside-like fragment ions [27, 28]. Fragment ions complementary to metal-coordinated $(w_3\text{-B})^-$, x_3 , x_4 , and z_4 are missing.

The localization of the binding sites of Zn^{2+} -ions

yields similar results as found for the Fe^{2+} -ions. Based on the observation of the position-defining metal-complexed $(w_3\text{-B}_3)^-$, $(w_3\text{-B}_4)^-$, and $[w_4\text{-d}_4]$ -fragment ions, the binding site of the Zn^{2+} -ions can be assigned to the three 3'-terminal phosphate groups and the two interjacent nucleosides (Figure 8).

Like the iron ions, the zinc ions are also located in the central region of the molecules, exerting a stabilizing effect on the phosphate backbone. There is evidence that the zinc ions coordinate to the phosphate groups. However, the exact binding sites cannot be appointed unambiguously. It stands to question whether nucleobases B_3 and B_4 of the interjacent nucleosides are involved in the metal coordination or not. The four $[\text{Zn}^{2+}]_3$ -pentamer complexes investigated show neither a base-specific, nor a positionally preferred loss of nucleobases B_3 and B_4 . Loss of B_3 from TTCAT is indicated by the $w_3\text{-C}$ fragment ion, and loss of B_4 from TTGCT and TTGAT results in the corresponding $w_3\text{-B}_4$ fragment ions. In the spectrum of $[\text{Zn}^{2+}]_3$ -TTGGT the $w_3\text{-G}$ ion is identified unambiguously. Consequently, no plausible pattern for the binding of Zn^{2+} -ions to the nucleobases B_3 and B_4 can be established. Such stochastic cleavages strongly indicate that the nucleobases are most unlikely to be involved in the coordination of the zinc ions. Also, in contrast to the observations made for Fe^{2+} -pentamer complexes, no $[\text{Zn}^{2+}]_5$ -pentamer complex was generated, indicating that zinc ions are not likely to coordinate to the nucleobases. Therefore, we propose the phosphate groups as the only binding sites for the zinc ions.

The difference between TTCAT/TTGAT and TTGCT/TTGGT, discussed for the bare oligonucleotides and for $[\text{Fe}^{2+}]_3$ -pentamer complexes, reappears for the $[\text{Zn}^{2+}]_3$ -pentamer complexes. Surprisingly, there is a clear difference between $[\text{Fe}^{2+}]_3$ - and $[\text{Zn}^{2+}]_3$ -pentamer complexes. While from the $[\text{Fe}^{2+}]_3$ -pentamer complexes mainly guanine and cytosine are cleaved off, giving abundant M-G- and M-C-ions, it is thymine and thymine in combination with water which are released from the Zn^{2+} -complexed oligonucleotides. M-G- and M-C-ions usually appear to a lesser extent only, except for the quite abundant M-C-ion in the spectrum of TTGCT. The observed preference for the loss of thymine may be attributed to the terminal position of the thymine bases versus the central position of guanine and cytosine. Despite its significantly lower proton affinity, protonation and subsequent loss of thymine to such an extent is somewhat unexpected.

Comparing the results from $[\text{Zn}^{2+}]_3$ -pentamer complexes with the data from $[\text{Fe}^{2+}]_3$ -pentamer complexes reveals the different influences of the two doubly charged metal ions on the pentanucleotides.

Conclusions

Results disclose a set of fundamental aspects of the gas-phase behavior of metal-oligonucleotide complexes. The identity of the metal ions and the base

sequences of the oligonucleotides both noticeably direct the fragmentation of the metal-oligonucleotide complexes upon CID. Assignment of metal-coordinated fragment ions and the appearance of terminal metal-free fragment ions provide new insight into the binding patterns of the three selected transition metals.

Fe^{3+} -complexes:

The coordination of a single Fe^{3+} -ion to a pentanucleotide results in a significant change of the fragmentation of the newly formed complex. The binding site of the Fe^{3+} -ion can unambiguously be appointed to the second phosphate group from the 5'-end of all four pentanucleotides investigated. Involvement of the nucleobases in the coordination of the metal ion is clearly excluded.

Fe^{2+} -complexes:

Investigation of $[Fe^{2+}]_3$ -pentanucleotide complexes exclusively reveals metal-complexed 3'-terminal fragment ions and the three Fe^{2+} -ions can be located at the three 3'-terminal phosphate groups. On the basis of additional experiments on $[Fe^{2+}]_5$ -pentanucleotide complexes it can be assumed that Fe^{2+} -ions predominantly occupy phosphate groups of the oligonucleotide backbone. If no further phosphates are available and additional binding sites are required, the nucleobases can serve as binding sites as well.

Zn^{2+} -complexes:

Experiments revealed the phosphate groups of the oligonucleotide backbone as the binding sites for Zn^{2+} -ions. Unlike for iron complexes, participation of the nucleobases in the binding of zinc ions in $[Zn^{2+}]_3$ -complexes cannot be excluded unambiguously.

This study demonstrates the potential of tandem mass spectrometry for characterizing noncovalent interactions between oligonucleotides and metal ions. Once the fundamental aspects of gas-phase dissociation of metal-oligonucleotide complexes are better understood, it should prove possible to characterize the binding motifs and binding stoichiometries of metal ions, and consequently, to gain deeper insight into the concepts of metal ion coordination to oligonucleotides.

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